



NOTESPK TEST SERIES

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Name:		Roll #:		Subject:	Chemistry-11	Test #:	337364
Test Detail:	Type 1B - 100 MCQs Test - Marks=100					Time:	
Syllabus:	Full Book					Date:	25-Nov-202

1- Circle the correct one.

(100 x 1=100)

- What is oxidation number of sulfur in SO_3 ? (A) (B) (C) (D)
- What is oxidation number of Si in SiCl_4 ? (A) (B) (C) (D)
- Select the two normal elements present in sixth period: (A) K, Ca (B) Rb, Sr (C) Cs, Ba (D) La, Hf
- What does the group number represent? (A) (B) (C) (D)
- Mark the correct statement: (A) Na^+ is smaller than Na atom
(B) Na^+ is larger than Na atom
(C) Cl^- is smaller than Cl atom
(D) Cl^- ion and Cl atom are equal in size
- What happens when acidic chlorides react with water? (A) (B) (C) (D)
- The hydrides of Group IA are: (A) Ionic (B) Covalent (C) Metallic
(D) Interstitial
- The number of neutrons presents in $^{39}_{19}\text{K}$ is: (A) 39 (B) 19 (C) 20 (D) 21
- What is the formula to find number of neutrons (N)? (A) (B) (C) (D)
- Which particle is deflected the most in a magnetic field? (A) (B) (C) (D)
- What type of spectrum shows coloured lines emitted by hot gases? (A) (B) (C) (D)
- What is the range of values for azimuthal quantum number (A) (B) (C) (D)
- The three p orbitals in a given energy level are oriented: (A) Along the same axis.
(B) At 45° to each other.
(C) Mutually perpendicular to each other along the x, y, and
(D) In a complex tetrahedral arrangement.
- Which sum determines the filling order of subshells? (A) (B) (C) (D)
- When 6d orbital is complete, the entering electron goes into: (A) 7f (B) 7s (C) 7p (D) 7d
- The state of hybridization in ethene molecule is: (A) dsp^2 (B) sp^3 (C) sp^2 (D) sp
- What is hybridization? (A) (B) (C) (D)
- What type of bonds occupy the corners in SO_3 ? (A) (B) (C) (D)
- What does a receptor mostly bind to? (A) (B) (C) (D)
- Bond order of helium molecule is: (A) Two (B) Three (C) Zero (D) One
- The bond order of N_2 molecule according to Molecular Orbital Theory is: (A) Zero (B) 01 (C) 02 (D) 03

22. Which of the following species has unpaired electrons in antibonding molecular orbitals? (A) O_2^{2+} (B) N_2^{2-} (C) B (D) F_2
23. What does balanced chemical equation have equal on both sides? (A) (B) (C) (D)
24. The number of moles of CO_2 which contain 8.0 of oxygen: (A) 0.25 (B) 0.15 (C) 0.35 (D) 1.45
25. A gas has a density of 1.96 g dm^{-3} at STP (0°C and 1.00 atm). What is its molecular mass? (Molar volume at STP = $22.4 \text{ dm}^3 \text{ mol}^{-1}$) (A) 11.2 g mol^{-1} (B) 22.4 g mol^{-1} (C) 44.0 g mol^{-1} (D) 88.0 g mol^{-1}
26. Which Formula is used to find volume at: (A) (B) (C) (D)
27. Which formula is used to find the number of moles from volume and concentration? (A) (B) (C) (D)
28. How much mass of NH_3 is formed from $28.0 \text{ g } N_2$ and $6 \text{ g } H_2$? (A) (B) (C) (D)
29. Where are steroids detected in athletes? (A) (B) (C) (D)
30. What do liquid crystal sensors measure? (A) (B) (C) (D)
31. Why is particle volume negligible in ideal gas? (A) (B) (C) (D)
32. When happens to volume of gas when pressure doubles? (A) (B) (C) (D)
33. What is surface tension of acetone at 20°C ? (A) (B) (C) (D)
34. What causes cooling during evaporation? (A) (B) (C) (D)
35. Vapour pressure of liquid depends upon: (A) Amount of liquid (B) Surface area (C) Temperature (D) Size of container
36. Glycerine boils at 210°C when external pressure is: (A) 760 torr (B) 500 torr (C) 100 torr (D) 50 torr
37. What type of reaction absorbs heat from surrounding? (A) (B) (C) (D)
38. Which Law is applied to Born-Haber cycle? (A) (B) (C) (D)
39. What is bond dissociation energy? (A) (B) (C) (D)
40. Which of the following reactions has an enthalpy change that is equal to the standard enthalpy of formation of water, $\Delta H_f^\circ[H_2O(l)]$? (A) $2H(g) + O(g) \rightarrow H_2O(l)$ (B) $H_2(g) + \frac{1}{2}O_2(g) \rightarrow H_2O(g)$ (C) $H_2(g) + \frac{1}{2}O_2(g) \rightarrow H_2O(l)$ (D) $2H^+_{(aq)} + O^{2-}_{(aq)} \rightarrow H_2O(l)$
41. What does a calorimeter measure? (A) (B) (C) (D)
42. What is released when food is digested? (A) (B) (C) (D)
43. What is bond energy of H-H used? (A) (B) (C) (D)
44. What is studied in reaction kinetics? (A) (B) (C) (D)
45. The rate of reaction determined at any given time is called: (A) Average rate (B) Instantaneous rate (C) Spontaneous rate (D) Overall rate
46. How is rate of reaction defined? (A) (B) (C) (D)
47. The process in which catalyst and reactant are in different phases is called: (A) Homogeneous catalysis (B) Heterogeneous catalysis (C) Auto catalysis (D) Negative catalysis
48. The catalyst which helps manufacture HNO_3 from ammonia oxidation? (A) (B) (C) (D)

49. Who gives the order of reaction? (A) (B) (C) (D)
50. Which symbol represents the rate constant? (A) (B) (C) (D)
51. What does the position of equilibrium refer to? (A) (B) (C) (D)
52. What occurs if is removed at equilibrium? (A) (B) (C) (D)
53. Which side does the reaction shift when volume decreases? (A) (B) (C) (D)
54. Which principle predicts direction of reaction with change in temperature? (A) (B) (C) (D)
55. What starts with maximum rate & then slows down in forward reactions? (A) (B) (C) (D)
56. What role does a catalyst play in equilibrium? (A) (B) (C) (D)
57. For which of the following reaction, the unit of equilibrium constant (K_c) is reciprocal of molar concentration (M^{-1}):
 (A) $3H_2(g) + N_2(g) \rightleftharpoons 2H_3N(g)$ (B) $2NO_2(g) \rightleftharpoons N_2O_4(g)$
 (C) $H_2(g) + I_2(g) \rightleftharpoons 2HI(g)$ (D) $N_2(g) + O_2(g) \rightleftharpoons 2HO(g)$
58. Who accepts protons in acid base indicator? (A) (B) (C) (D)
59. The ionization constant of pure water at 25°C is:
 (A) $1.8 \times 10^{-16} \text{ moles dm}^{-3}$ (B) $1.6 \times 10^{-16} \text{ moles dm}^{-3}$
 (C) $1.0 \times 10^{-14} \text{ moles}^2 \text{ dm}^{-6}$
 (D) $1.8 \times 10^{-14} \text{ moles}^2 \text{ dm}^{-6}$
60. The pH of $10^{-4} \text{ moles / dm}^3$ of $Ba(OH)_2$ is:
 (A) 4.5 (B) 6.4 (C) 7.5 (D) 10.3
61. What is the formula used to find $[H_3O^+]$ from K_a and concentration? (A) (B) (C) (D)
62. By adding NH_4Cl to NH_4OH solution. The ionization of NH_4OH :
 (A) Increases (B) Remains same (C) Decreases
 (D) Increases 100 times
63. Mixture of NH_4OH and NH_4Cl is one of the best example of:
 (A) Acidic buffer (B) Basic buffer
 (C) Common ion effect (D) Solubility product
64. What is K_{sp} known as? (A) (B) (C) (D)
65. What system is used as reference for electrode potential? (A) (B) (C) (D)
66. Which of the following half-reactions represents the oxidation process occurring in the disproportionation of Cu^+ ?
 (A) $Cu_{(aq)}^{2+} + 2e^- \rightarrow Cu_{(s)}$ (B) $Cu_{(aq)}^+ + e^- \rightarrow Cu_{(s)}$
 (C) $Cu_{(aq)}^+ \rightarrow Cu_{(aq)}^{2+} + e^-$ (D) $2Cu_{(aq)}^+ \rightarrow Cu_{(aq)}^{2+} + Cu_{(s)}$
67. Which group has fixed oxidation number + 1? (A) (B) (C) (D)
68. What is oxidation number of oxygen in H_2O_2 ? (A) (B) (C) (D)
69. Who acts as both oxidizing and reducing agent simultaneously? (A) (B) (C) (D)
70. Which electrode is positively charged in an electrolytic cell? (A) (B) (C) (D)
71. What gas is formed at the anode when chloride ions are oxidized? (A) (B) (C) (D)
72. Which hydrocarbons have high carbon to hydrogen ratio? (A) (B) (C) (D)
73. What do you do when two substituents are on the same carbon? (A) (B) (C) (D)
74. What type of bond cleavage produces ions? (A) (B) (C) (D)

75. The catalytic oxidation of methane produces: (A) $CO + H_2O$ (B) $CO_2 + H_2O$ (C) $C_2 + H_2O$
(D) $H_3C - OH$
76. $CH_4 + Cl_2 \xrightarrow[\text{Sunlight}]{\text{Diffused}}$ product is: (A) Chloroform only (B) Carbon tetra chloride
(C) $CH_3Cl + CH_2Cl_2$ (D) Mixture of A, B, C
77. What intermediate is formed during the electrophilic addition of HBr to an alkene? (A) Carbocation (B) Carbanion (C) Radical
(D) Epoxide
78. Which of the following carbocations would be the least stable? (A) $(CH_3)_2CH^+$ (B) $CH_2CH_2^+$ (C) CH_3^+
(D) $(CH_3)_3C^+$
79. Despite being the most abundant gas in the Earth's atmosphere, nitrogen does not readily participate in combustion reactions because: (A) It is denser than oxygen.
(B) It has a high specific heat capacity.
(C) Breaking the $N \equiv N$ bond requires a large amount of energy.
(D) It is a noble gas.
80. What are components of gun powder? (A) (B) (C) (D)
81. Which Nitrogen oxide is reddish brown gas? (A) (B) (C) (D)
82. Which of the elements gives acidic oxide? (A) N (B) As (C) Sb (D) Bi
83. What does a catalytic converter reduce? (A) (B) (C) (D)
84. What is released during denitrification, (A) (B) (C) (D)
85. Which sulfur state has six unpaired electrons? (A) (B) (C) (D)
86. Which halogen reacts readily with water? (A) (B) (C) (D)
87. Which halogens are very rare and radioactive? (A) (B) (C) (D)
88. Water is disinfected by a substance to avoid toxification: (A) $KMnO_4$ (B) Alums (C) O_3 (D) Cl_2
89. Why is chlorine more volatile than bromine and iodine? (A) (B) (C) (D)
90. Which forces affect halogen volatility? (A) (B) (C) (D)
91. Who has the weakest bond strength among halogens? (A) (B) (C) (D)
92. Why is fluorine the most reactive halogen? (A) Bond length in the halogen molecule
(B) Bond strength in the halogen molecule
(C) Electronegativity of the halogen
(D) Number of electrons in the halogen molecule
93. How does halogen reactivity change down the group? (A) (B) (C) (D)
94. Who are affected by the impact of burning fossil fuels? (A) (B) (C) (D)
95. What is the one effect of human activities on the atmosphere? (A) (B) (C) (D)
96. Which is a secondary pollutant? (A) Carbonic acid (B) CO_2 (C) SO_2
(D) CO
97. Which natural source produces hydrogen sulphide? (A) (B) (C) (D)
98. What is the major cause of acid rain? (A) (B) (C) (D)
99. What is the primary cause of climate change? (A) (B) (C) (D)
100. The pH of unpolluted rain water should be: (A) 5.00 (B) 5.60 (C) 6.50 (D) 7.00

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TEST TYPE WITH ANSWERS KEY

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20
A	C	C	B	A	B	A	C	B	C	B	A	C	C	C	C	A	B	A	C
21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36	37	38	39	40
D	B	A	A	C	D	A	D	D	D	A	B	A	B	C	D	B	D	B	C
41	42	43	44	45	46	47	48	49	50	51	52	53	54	55	56	57	58	59	60
A	A	D	A	B	D	B	C	B	C	D	C	A	B	B	B	C	B	C	D
61	62	63	64	65	66	67	68	69	70	71	72	73	74	75	76	77	78	79	80
C	C	B	A	D	C	B	B	B	C	D	C	C	B	D	A	C	C	A	A
81	82	83	84	85	86	87	88	89	90	91	92	93	94	95	96	97	98	99	100
A	A	D	B	C	A	C	C	A	D	C	A	B	D	D	A	C	C	B	B

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